

## PHYS 111-01 STARTING EQUATIONS - WAVES, SOUND, THERMAL PHYSICS

The equations given below may be used as starting points for solving problems on the final.

### WAVES AND SOUND

#### Waves:

$$v_w = \lambda f \quad f = (1/T) = \omega/2\pi$$

$$v_w = \sqrt{\frac{F}{m/L}} \quad (\text{stretched string})$$

#### Sound:

$$f_n = (n/2L)v \quad \text{normal modes - stretched string or open tube}$$

$$f_n = (n/4L)v \quad \text{normal modes - tube with only one end open}$$

$$I = P/A \quad I(r) = P/4\pi r^2 \quad (\text{spherical sound wave})$$

#### Doppler Shift – Easy versions

$f' = f[v_w/(v_w \mp u_s)]$  moving source  $f' = f[(v_w \pm u_o)/v_w]$  moving observer (Top signs, approaching)

#### Doppler Shift – Textbook versions

$$f' = \left(1 \pm \frac{u_o}{v_w}\right) f \quad \text{Moving Observer} \quad f' = \left(\frac{1}{1 \mp \frac{u_s}{v_w}}\right) f \quad \text{Moving Source (Top signs, approaching)}$$

### TEMPERATURE AND HEAT

$$\Delta L = \alpha L_0 \Delta T \quad \Delta V = \beta V_0 \Delta T \quad Q = mc\Delta T \quad Q = mL$$

#### Ideal Gas

$$PV = NkT = nRT$$

$$U = (3/2) NkT = (3/2)nRT \quad \text{Monatomic Ideal Gas}$$

#### Thermodynamics, Thermal Conduction, Thermal Radiation:

$$Q = kA(\Delta T/L)$$

$$P = eA\sigma T^4$$

$$\Delta U = Q_{in} - W_{by}$$

$$W = P\Delta V$$

$$\Delta S = Q/T$$

$$\Delta S_{universe} \geq 0$$

#### Heat Engines

$$e = W/Q_h \quad e_{max} = 1 - (T_c/T_h)$$

#### QUANTITY

#### SYMBOL

#### UNIT OR VALUE

#### COMMENTS

Period

T

s

Frequency

f

Hz

Hz = sec<sup>-1</sup>

Wavelength

$\lambda$

m

Wave velocity

v or  $v_w$

m/s

$v_w = 343$  m/s, sound in 20°C air

Intensity

I

W/m<sup>2</sup>

Temperature

T

K or °C

Linear expansion coefficient  $\alpha$

K<sup>-1</sup> or °C<sup>-1</sup>

Volume expansion coefficient  $\beta$

K<sup>-1</sup> or °C<sup>-1</sup>

Specific Heat

c

J/(kg-K) or J/(kg-°C)

Heat

Q

J

Latent Heat

L

J/kg

Thermal Conductivity

k

J/(s-m-C°)

Stefan-Boltzmann Constant

$\sigma$

$5.67 \times 10^{-8}$  W/(m<sup>2</sup>-K<sup>4</sup>)

Boltzmann Constant

k

$1.38 \times 10^{-23}$  J/K

Gas Constant

R

8.31 J/(mol-K)

Number of moles

n

Entropy

S

J/K

Internal Energy

U

J